

Quiz 4 – September 14, 2018

Complete the following problems. Write your final answers in the blanks provided. You must show your work to receive full credit. Show your answers to the correct number of significant figures with the correct units.

1. Malachite, a copper-containing mineral with the formula $\text{Cu}_2(\text{OH})_2\text{CO}_3$ can be used to prepare copper metal. Answer the following regarding malachite.

a. What is the molar mass of malachite? (2 pts.)

From the formula, 1 mol malachite consists of 2 mol Cu, 5 mol O, 2 mol H and 1 mol C, so the molar mass is:
 $[2(63.546) + 5(15.999) + 2(1.008) + 1(12.011)]\text{g/mol} = 221.114 \text{ g/mol}$

Answer 221.114 g/mol

b. Determine the mass percent of oxygen in malachite. $\text{Cu}_2(\text{OH})_2\text{CO}_3$? (3 pts.)

Molar mass of malachite is 221.116 g/mol and each mole of malachite contains 5 moles of oxygen, or
 $5 \times 15.9994 \text{ g} = 79.997 \text{ g oxygen}$, so the mass % oxygen is:

$$\frac{79.995 \text{ g O}}{221.114 \text{ g malachite}} \times 100\% = 36.18 \% \text{ O}$$

Answer 36.18 % O

c. If you wanted to prepare 10.0 grams of copper from malachite, what mass of the mineral would be required? (3 points)

Each mole of malachite contains 2 mol Cu, so 221.114 g malachite contains $2(63.546) = 127.092 \text{ g Cu}$

$$10.0 \text{ g Cu} \times \frac{221.114 \text{ g malachite}}{127.092 \text{ g Cu}} = 17.398 \text{ g malachite}$$

Answer 17.4 g malachite

2. Complete the table of ions below: (8 points)

Name	Formula (including charge)
hydroxide ion	OH^-
ammonium ion	NH_4^+
magnesium ion	Mg^{2+}
carbonate ion	CO_3^{2-}

Name	Formula (including charge)
acetate ion	$\text{C}_2\text{H}_3\text{O}_2^-$
perchlorate ion	ClO_4^-
sulfur anion (aka sulfide)	S^{2-}
cyanide ion	CN^-

3. Fructose is a simple sugar that is found in many plants as a component of sucrose. Excessive consumption of fructose has been linked to insulin dependence, type 2 diabetes, and cardiovascular disease. If fructose has a percent composition of 40.00% C, 6.71% H and 53.28% O and a molar mass of 180.16 g/mol, what are the empirical and molecular formulas of fructose? (9 pts.)

There are at least two approaches to this problem. I'll show you two, either approach is fine.

Approach 1:

Let's assume we have 100 g of fructose. Then we will have 40.00g C, 6.71g H and 52.28g O. How many moles of each?

$$40.00 \text{ g-C} \times \frac{1 \text{ mol C}}{12.011 \text{ g-C}} = 3.33 \text{ mol C}$$

$$6.71 \text{ g-H} \times \frac{1 \text{ mol H}}{1.008 \text{ g-H}} = 6.66 \text{ mol H}$$

Similarly, we find we have 3.33 mol O.

Therefore, our empirical formula is $C_{3.33}H_{6.66}O_{3.33}$, or $C_1H_2O_1$. If this were also the molecular formula, the formula weight would be equal to the molecular mass. For the empirical formula, the formula weight is 30.026 g/mol, which is one sixth ($180.16/30.026 = 6.00$) of the molecular mass, therefore, the molecular formula must be $C_6H_{12}O_6$.

Approach 2:

We know that one mole of fructose has a mass of 180.16 grams and we know the percent composition of fructose, so we can calculate how many moles of each atom must be present in one mole of fructose.

$$\frac{40.00 \text{ g-C}}{100 \text{ g-fructose}} \times \frac{180.16 \text{ g-fructose}}{1 \text{ mol fructose}} \times \frac{1 \text{ mol C}}{12.011 \text{ g-C}} = \frac{6.00 \text{ mol C}}{\text{mol fructose}}$$

$$\frac{6.71 \text{ g-H}}{100 \text{ g-fructose}} \times \frac{180.16 \text{ g-fructose}}{1 \text{ mol fructose}} \times \frac{1 \text{ mol H}}{1.008 \text{ g-H}} = \frac{12.00 \text{ mol H}}{\text{mol fructose}}$$

$$\frac{53.28 \text{ g-O}}{100 \text{ g-fructose}} \times \frac{180.16 \text{ g-fructose}}{1 \text{ mol fructose}} \times \frac{1 \text{ mol O}}{15.999 \text{ g-O}} = \frac{6.00 \text{ mol O}}{\text{mol fructose}}$$

Therefore, our molecular formula must be $C_6H_{12}O_6$ and the empirical formula is CH_2O .

Answer: Empirical Formula CH₂O **Molecular Formula** C₆H₁₂O₆

Periodic Table of the Elements

IA 1A												VIII 8A					
1 H Hydrogen 1.008	2 He Helium 4.003											13 B Boron 10.811	14 C Carbon 12.011	15 N Nitrogen 14.007	16 O Oxygen 15.999	17 F Fluorine 18.998	18 Ne Neon 20.180
3 Li Lithium 6.941	4 Be Beryllium 9.012											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.305	3 III 3B	4 IV 4B	5 V 5B	6 VI 6B	7 VII 7B	8 VIII 8	9 VIII 8	10 VIII 8	11 IB 1B	12 IIB 2B	13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.631	33 As Arsenic 74.922	34 Se Selenium 78.971	35 Br Bromine 79.904	36 Kr Krypton 83.798
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.414	49 In Indium 114.818	50 Sn Tin 118.711	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.904	54 Xe Xenon 131.294
55 Cs Cesium 132.905	56 Ba Barium 137.328	57-71	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.085	79 Au Gold 196.967	80 Hg Mercury 200.592	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [208.982]	85 At Astatine 209.987	86 Rn Radon 222.018
87 Fr Francium 223.020	88 Ra Radium 226.025	89-103	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [278]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [280]	112 Cn Copernicium [285]	113 Nh Nihonium [286]	114 Fl Flerovium [289]	115 Mc Moscovium [289]	116 Lv Livermorium [293]	117 Ts Tennessine [294]	118 Og Oganesson [294]

Lanthanide Series	57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.243	61 Pm Promethium 144.913	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.055	71 Lu Lutetium 174.967
Actinide Series	89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.064	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.070	98 Cf Californium 251.080	99 Es Einsteinium [254]	100 Fm Fermium 257.095	101 Md Mendelevium 258.1	102 No Nobelium 259.101	103 Lr Lawrencium [262]