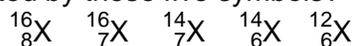Complete each of the following problems. For numerical problems, you must show your work in order to possibly earn full credit.

1. Consider the isotope symbols below, where X replaces each element symbol. How many different **<u>elements</u>** are represented by these five symbols? Identify each element. (3 pts.)

$$^{16}_{8}X \quad ^{16}_{7}X \quad ^{14}_{7}X \quad ^{14}_{6}X \quad ^{12}_{6}X$$

The identity of an element is determined by its atomic number, which is given in the subscript of each symbol. Therefore, there are three elements shown, atomic numbers 6, 7, and 8. These elements are carbon, nitrogen and oxygen.

2. Hydrogen gas can be made by decomposing methane (CH_4), the main component in natural gas. If the mass ratio for methane is 3.00 grams carbon (C) to 1.00 gram hydrogen (H), how many grams of hydrogen can be made from 90.0 grams of methane? (3 pts.)

Since the compound has 3 grams of C per 1 gram H, the mass of methane is one fourth hydrogen:

$$\frac{1.00 \text{ g hydrogen}}{(3.00 \text{ g carbon} + 1.00 \text{ g hydrogen})} = \frac{1.00 \text{ g hydrogen}}{4.00 \text{ g methane}}$$

With this ratio, we can determine the mass of hydrogen in 90.0 g methane:

$$\frac{1.00 \text{ g hydrogen}}{4.00 \text{ g methane}} \times 90.0 \text{ g methane} = 22.5 \text{ g hydrogen}$$

3. A cube of gold (Au) that is 1.00 cm on each side has a mass of 19.32 g. How many gold atoms does this cube contain? (4 pts.)

This problem requires us to convert the mass of gold to a number of atoms. Avogadro's number will allow the conversion of *moles* to atoms, but first we need to convert to moles using the molar mass.

$$19.32 \text{ g Au} \times \frac{1 \text{ mol Au}}{196.967 \text{ g Au}} = 0.09809 \text{ moles Au}$$

Now Avogadro's number will allow us to calculate atoms:

$$0.09809 \text{ mol Au} \times \frac{6.022 \times 10^{23} \text{ atoms Au}}{1 \text{ mol Au}} = 5.91 \times 10^{22} \text{ atoms Au}$$

Possibly Useful Information

$N_A = 6.022 \times 10^{23}$ anythings per mole	density = mass/volume

To save some calculation time, you may round all atomic masses to two (2) decimal points.

Periodic Table

1 / 1A	2 / 2A	3 / 3B	4 / 4B	5 / 5B	6 / 6B	7 / 7B	8 / 8B	9 / 8B	10 / 8B	11 / 1B	12 / 2B	13 / 3A	14 / 4A	15 / 5A	16 / 6A	17 / 7A	18 / 8A
1 H 1.00794																	2 He 4.00260
3 Li 6.941	4 Be 9.01218											5 B 10.811	6 C 12.011	7 N 14.0067	8 O 15.9994	9 F 18.9984	10 Ne 20.1797
11 Na 22.9898	12 Mg 24.3050											13 Al 26.9815	14 Si 28.0855	15 P 30.9738	16 S 32.066	17 Cl 35.4527	18 Ar 39.948
19 K 39.0983	20 Ca 40.078	21 Sc 44.9559	22 Ti 47.88	23 V 50.9415	24 Cr 51.9961	25 Mn 54.9381	26 Fe 55.847	27 Co 58.9332	28 Ni 58.693	29 Cu 63.546	30 Zn 65.39	31 Ga 69.723	32 Ge 72.61	33 As 74.9216	34 Se 78.96	35 Br 79.904	36 Kr 83.80
37 Rb 85.4678	38 Sr 87.62	39 Y 88.9059	40 Zr 91.224	41 Nb 92.9064	42 Mo 95.94	43 Tc (98)	44 Ru 101.07	45 Rh 102.906	46 Pd 106.42	47 Ag 107.868	48 Cd 112.411	49 In 114.818	50 Sn 118.710	51 Sb 121.757	52 Te 127.60	53 I 126.904	54 Xe 131.29
55 Cs 132.905	56 Ba 137.327	57 *La 138.906	72 Hf 178.49	73 Ta 180.948	74 W 183.84	75 Re 186.207	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.967	80 Hg 200.59	81 Tl 204.383	82 Pb 207.2	83 Bi 208.980	84 Po (209)	85 At (210)	86 Rn (222)
87 Fr (223)	88 Ra 226.025	89 †Ac 227.028	104 Rf (261)	105 Db (262)	106 Sg (266)	107 Bh (264)	108 Hs (277)	109 Mt (268)	110 Ds (271)	111 Rg (272)							

*Lanthanide series

58 Ce 140.115	59 Pr 140.908	60 Nd 144.24	61 Pm (145)	62 Sm 150.36	63 Eu 151.965	64 Gd 157.25	65 Tb 158.925	66 Dy 162.50	67 Ho 164.930	68 Er 167.26	69 Tm 168.934	70 Yb 173.04	71 Lu 174.967

†Actinide series

90 Th 232.038	91 Pa 231.036	92 U 238.029	93 Np 237.048	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)