Chem 130 **Name**______________________________

Exam 3, Ch 14 and 15 **December 7, 2011**

100 Points

Please follow the instructions for each section of the exam. Show your work on all mathematical problems. Provide answers with the correct units and significant figures. Be concise in your answers to discussion questions.

Part 0: Warmup. 4 points each

1. Half-life is independent of initial concentration for ________________ reactions :
   a. zero order
   b. first order
   c. second order
   d. no
   e. all

   Answer ____________

2. The pressure of a reaction vessel that contains an equilibrium mixture in the reaction $SO_2Cl_2(g) \rightleftarrows SO_2(g) + Cl_2(g)$ is increased. When equilibrium is reestablished:
   a. the amount of Cl_2 will have increased.
   b. the amount of SO_2 will have decreased.
   c. the amounts of SO_2 and Cl_2 will have remained the same.
   d. the amounts of all SO_2, Cl_2, and SO_2Cl_2 will have decreased..

   Answer ____________

3. For the reaction $CO(g) + 2H_2(g) \rightleftarrows CH_3OH(g)$ $K_c = 14.5$. If 5.00 mol CO, 2.00 mol H_2 and 3.00 mol of CH_3OH are brought together and allowed to react, which of the following describes the composition of the system at equilibrium?
   a. Some CO and H_2 will have been consumed to make more CH_3OH.
   b. Some CH_3OH will have been consumed to make more CO and H_2
   c. The amounts of CO, H_2, and CH_3OH will be unchanged from their initial values.
   d. There is not enough information to determine the equilibrium composition.

   Answer ____________

Part I: Complete problem 4. (10 points)

4. The decomposition of hydrogen peroxide (H_2O_2) to liquid water and oxygen gas is a <u>spontaneous</u> process, but occurs <u>slowly</u>, allowing for a reasonable shelf-life for peroxide solutions. In the presence of a catalyst (like the iron in your blood), the decomposition is much <u>more rapid</u>. Draw and correctly label reaction coordinate diagrams that represent each of the two situations and describe how each diagram reflects the thermodynamics and kinetics of the situation.

Part II Kinetics. Answer <u>three (3)</u> of problems 5-8. Clearly mark the problems you do not want graded. 15 points each. (13 points each)

Use the tabulated data below for problems 5 and 6. The data corresponds to the reaction below and shows dependence of concentration of A on time in experiments run at three temperatures.

$$2A \rightarrow B + C$$

	T = 298 K	T = 348 K	T = 398 K
Time (sec)	[A] (M)	[A] (M)	[A] (M)
0	0.1000	0.1000	0.1000
1	0.0999	0.0991	0.0940
2	0.0999	0.0983	0.0883
8	0.0995	0.0932	0.0609
32	0.0980	0.0755	0.0137

5. Consider the data above. Clearly describe how the data could be used to determine the order for the reaction in A. No calculations are necessary. Feel free to include appropriate diagrams.

6. Consider the data above. Assuming you know the reaction to be first order in A, describe how you could determine the activation energy for the reaction. No calculations are necessary. Feel free to include appropriate diagrams.

7. The reaction $I^-(aq) + OCl^-(aq) \rightarrow IO^-(aq) + Cl^-(aq)$ was studied and the data below were obtained. Determine the rate law and the value of the rate constant for this reaction.

$[I^-]_0$(mol/L)	$[OCl^-]_0$(mol/L)	Initial Rate (mol/Ls)
0.12	0.18	0.0791
0.060	0.18	0.0395
0.030	0.090	0.00988
0.24	0.090	0.0791

8. Experiment has shown that the rate law for the reaction $2NO(g) + Cl_2(g) \rightarrow 2NOCl(g)$ is Rate = $k[NO][Cl_2]$. One proposed mechanism for this process is shown below, with the second step being rate-determining. Is this a reasonable mechanism for the reaction? Justify your decision.

$$NO + Cl_2 \underset{k_{-1}}{\overset{k_1}{\rightleftarrows}} NOCl_2 \quad \text{(fast)}$$

$$NOCl_2 + NO \xrightarrow{k_2} 2NOCl \quad \text{(slow)}$$

Part III Equilibrium. Answer <u>three (3)</u> of problems 9-12. Clearly mark the problems you do not want graded. (13 points each)

9. What do we mean when we say a system has *come to equilibrium*? Describe the equilibrium condition and why we don't use a single headed arrow when we write equilibria. What does a small equilibrium constant mean in terms of thermodynamics?

10. Consider the reaction below. If the initial concentrations of H_2, F_2, and HF are 0.0100M, 1.25 M, and 2.21 M, respectively, is the system at equilibrium? If not, which way will the reaction go to achieve the equilibrium condition? Set up, but do not complete the calculation you would use to determine the equilibrium concentrations of each of the species in the reaction. You DO NOT need to arrive at a numerical answer, just get to the point where you have one algebraic expression you could solve, given additional time. Be sure to tell me what you would do with the result of your calculation.

$$H_2(g) + F_2(g) \rightleftarrows 2HF(g) \qquad K = 115$$

11. Suppose the reaction system below has already reached equilibrium. Predict the effect of the following changes on the system. Justify your predictions with a brief statement.

$$UO_2(s) + 4HF(g) \rightleftarrows UF_4(g) + 2H_2O(g)$$

a. More UO_2 is added to the system.

b. The reaction is performed in a glass reaction vessel and the HF reacts with the glass.

c. Water vapor is removed.

d. The volume is increased.

12. You have been tasked with determining the equilibrium constant for the reaction of H_2 and S_2 gases to produce hydrogen sulfide. A mixture of 1.00 g H_2 and 1.00 g H_2S in a 0.500 L flask comes to equilibrium at 1670 K. At equilibrium, there is 8.00×10^{-6} mol of S_2 present. What are the values for K_c and K_p at this temperature?

Possibly Useful Information

$\text{slope} = m = \frac{\Delta y}{\Delta x} = \frac{y_2 - y_1}{x_2 - x_1}$	$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$	R = 0.08206 L atm mol^{-1} K^{-1} R = 8.314 J mol^{-1} K^{-1}
pV = nRT	$\Delta G = -RT\ln K$	$K_p = K_c(RT)^{\Delta n}$
$k = Ae^{-E_a/RT}$	$\ln k = -\left(\frac{E_a}{R}\right)\left(\frac{1}{T}\right) + \ln A$	$\ln\frac{k_2}{k_1} = \frac{E_a}{R}\left(\frac{1}{T_1} - \frac{1}{T_2}\right)$
rate = $k[A]^1$	$[A]_t = -kt + [A]_0$	$t_{1/2} = [A]_0/2k$
rate = $k[A]^2$	$\ln[A]_t = -kt + \ln[A]_0$	$t_{1/2} = 1/(k[A]_0)$
rate = $k[A]^0$	$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$	$t_{1/2} = 0.693/k$

1 1A	2 2A	3 3B	4 4B	5 5B	6 6B	7 7B	8 8B	9	10	11 1B	12 2B	13 3A	14 4A	15 5A	16 6A	17 7A	18 8A
1 H 1.00794																	2 He 4.00260
3 Li 6.941	4 Be 9.01218											5 B 10.811	6 C 12.011	7 N 14.0067	8 O 15.9994	9 F 18.9984	10 Ne 20.1797
11 Na 22.9898	12 Mg 24.3050											13 Al 26.9815	14 Si 28.0855	15 P 30.9738	16 S 32.066	17 Cl 35.4527	18 Ar 39.948
19 K 39.0983	20 Ca 40.078	21 Sc 44.9559	22 Ti 47.88	23 V 50.9415	24 Cr 51.9961	25 Mn 54.9381	26 Fe 55.847	27 Co 58.9332	28 Ni 58.693	29 Cu 63.546	30 Zn 65.39	31 Ga 69.723	32 Ge 72.61	33 As 74.9216	34 Se 78.96	35 Br 79.904	36 Kr 83.80
37 Rb 85.4678	38 Sr 87.62	39 Y 88.9059	40 Zr 91.224	41 Nb 92.9064	42 Mo 95.94	43 Tc (98)	44 Ru 101.07	45 Rh 102.906	46 Pd 106.42	47 Ag 107.868	48 Cd 112.411	49 In 114.818	50 Sn 118.710	51 Sb 121.757	52 Te 127.60	53 I 126.904	54 Xe 131.29
55 Cs 132.905	56 Ba 137.327	57 *La 138.906	72 Hf 178.49	73 Ta 180.948	74 W 183.84	75 Re 186.207	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.967	80 Hg 200.59	81 Tl 204.383	82 Pb 207.2	83 Bi 208.980	84 Po (209)	85 At (210)	86 Rn (222)
87 Fr (223)	88 Ra 226.025	89 †Ac 227.028	104 Rf (261)	105 Db (262)	106 Sg (266)	107 Bh (264)	108 Hs (277)	109 Mt (268)	110 Ds (271)	111 Rg (272)							

*Lanthanide series	58 Ce 140.115	59 Pr 140.908	60 Nd 144.24	61 Pm (145)	62 Sm 150.36	63 Eu 151.965	64 Gd 157.25	65 Tb 158.925	66 Dy 162.50	67 Ho 164.930	68 Er 167.26	69 Tm 168.934	70 Yb 173.04	71 Lu 174.967
†Actinide series	90 Th 232.038	91 Pa 231.036	92 U 238.029	93 Np 237.048	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)

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